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Mesoporous VO₂ nanowires with excellent cycling stability and enhanced rate capability for lithium batteries†

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To combine the merits of the one-dimensional structure and the porous structure, mesoporous VO₂ nanowires have been designed and reported for the first time. Excellent cycling stability and enhanced rate performance are obtained and may be attributed to the mesoporous nanowires, realizing both high surface area for more active sites and facile stress relaxation resulting in excellent structure stability. Our results demonstrate that the mesoporous nanowires are favourable for high-rate and long-life lithium batteries.

Introduction

Nowadays, the ever-increasing energy demand for new, sustainable and environmentally-friendly energy sources has been a vital and challenging issue.^{1–6} Among those various available energy storage technologies, rechargeable lithium batteries (LBs) have received great interest due to their high energy density, long lifespan and no memory effect.^{7–14} However, LBs fall short of the ever-increasing energy demands from large scale applications, such as electrical vehicles and grid-level energy storage. In order to meet the requirements of large scale applications, high power density is urgently required.^{15–20} Thus, it is of great significance to search for ideal electrode materials realizing high energy density and power density. Vanadium oxides are among the best cathode materials for rechargeable lithium batteries, due to both their large specific capacity and abundant sources. VO₂ (B) is one of most attractive electrode materials owing to the advantageous feature of double layers of V₄O₁₀ type for fast lithium intercalation/

deintercalation.^{21–24} However, the low diffusion coefficient of lithium ions (10^{–9} to 10^{–10} cm² s^{–1})^{25,26} and limited cycling stability^{27–31} hinder the practical applications of this material.

Designing electrode materials within nanoscale has been regarded as one of the most efficient method in fabricating ideal electrodes.^{32–34} Compared with bulk materials, nanowires are able to exhibit the advantages of both facile strain relaxation upon charging/discharging related to long cycle life and large surface to volume ratio to increase the contact area between the electrolyte and electrode. Nanowires as building blocks with complex structures and multiple functionalities have been reported to demonstrate the superior electrochemical performance in energy storage applications.^{35–40} In our previous work, constructing pores through top-down method in microplates has been demonstrated to be an efficient approach in fabricating electrode with excellent rate capability and stable capacity.⁴¹ Thus, constructing pores in nanowire will be a more efficient method because it can not only utilize superiorities of one-dimensional structure, but also take the advantages of mesoporous structure, such as good access of the electrolyte to the electrode surface, large surface area facilitating charge transfer across the electrode/electrolyte interface, increased utilization of active material and suppressed phase transformations and structure degradations during cycling.^{42–45} However, as is well known, nanowires due to the high reactivity are challenging to create pores without damaging the mechanical integrity. Thus, controllable and rational pores construction in nanowire has been a challenging issue.

Herein, to combine both merits of nanowire and mesoporous structure, a facile method preparing mesoporous VO₂ nanowires is reported *via* hydrothermal process and solid state reaction. NH₄V₃O₈ nanowires are obtained through the condensation of the vanadium precursor⁴⁵ *via* hydrothermal method. Through annealing treatment, the NH₄V₃O₈ nanowires are decomposed leading to internal gas emitting from inside and surface of the nanowire. The pressure of the internal gas may create interconnected pore networks, leading to the unique

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open pore distribution. In this way, the mesoporous VO₂ nanowires are achieved. Benefiting from the 1D structure and the tunable interconnected porosity, such unique architecture may greatly enhance the electrochemical performance in terms of specific capacity, cycling stability and rate capability.

Experimental section

Materials synthesis

In a typical synthesis, NH₄VO₃ (9 mmol) was dissolved in deionized water (40 mL) at 80 °C. The pH of the solution was adjusted to 2 by adding 5% hydrogen chloride solution. After stirring for two hours, the solution was then transferred to autoclave at 180 °C for 7 days. The precipitation was washed with water and alcohol for several times to get the NH₄V₃O₈ nanowires. Subsequently, the mesoporous VO₂ nanowires were obtained by annealing the as-prepared NH₄V₃O₈ nanowires in N₂ atmosphere at 400 °C for 4 h at the heating rate of 10 °C min⁻¹. The non-mesoporous VO₂ nanowires were obtained through oxalic acid reduction method. In a typical synthesis, the V₂O₅ powder (2 mmol) and oxalic acid (3.6 mmol) were dissolved in 33 mL water. The solution was then transferred into autoclave at 180 °C for 24 hours. The precipitation was then washed with water and alcohol for several times and dried at 60 °C for 8 h to obtain the non-mesoporous VO₂ nanowires.

Characterization

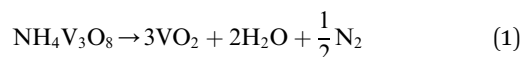
X-ray diffraction (XRD) measurements were performed to investigate the crystallographic information using a D8 Advance X-ray diffract meter with non-monochromatic Cu Kα X-ray 15 source. Field emission scanning electron microscopic (FESEM) images were collected with a JEOL JEM-7100F microscopy at an acceleration voltage of 10 kV. Transmission electron microscopic (TEM) images and high-resolution transmission electron microscopic (HRTEM) images were recorded by using a JEOL JEM-2100F FEF microscope. Brunauer–Emmet–Teller (BET) surface areas were measured using Tristar II 3020 instrument to investigate the adsorption of nitrogen. X-ray photoelectron spectroscopy (XPS) is carried out by using KRATOS XSAM800 electron spectrometer.

Measurement of electrochemical performance

The electrochemical properties were carried out by assembly of 2025 and 2016 coin cells in a glove box filled with pure argon gas. Lithium pellet was used as the anode; 1 M solution of LiPF₆ in ethylene carbon (EC)/dimethyl carbonate (DMC) as electrolyte. Cathode electrodes were obtained with 70% mesoporous VO₂ nanowires as active material, 20% acetylene black and 10% poly (tetrafluoroethylene) (PTFE). Galvanostatic charge-discharge cycling was studied in a potential range of 2.0–3.0 V vs. Li/Li⁺ with a multichannel battery testing system (LAND CT2001A). Cyclic voltammetry (CV) and AC-impedance spectra were tested with an electrochemical workstation (CHI 760D and Autolab PGSTAT 302).

Results and discussion

The schematic illustration of formation process is shown in Fig. 1. During the first process, the condensation of the vanadium precursor goes *via* oxalation as well as preferably *via* the faster olation reaction along the H₂O–V–OH direction and leads to the formation of 1D vanadium oxide within the xy plane.⁴⁶ During the second step, through different annealing temperature, the different valence of vanadium is obtained (Fig. S8†). The NH₄V₃O₈ nanowires were decomposed into mesoporous VO₂ nanowires under annealing process at annealing temperature of 450 °C *via* the reaction (1):



Owing to the gas emitting, the pores are interconnected and the nanowire is composed of many small pores. The vanadium oxide which blocks the gas from generating is induced by H₂O and N₂, leading to the pores both on the surface and the inner side of the nanowires. The interconnected porosity can lead to high surface area and, in this way, is able to offer more active sites and shorten the lithium diffusion distance for lithium insertion/deinsertion (Fig. 1B). Moreover, the porous structure may create void for facile stress relaxation in lithium insertion/deinsertion.

The products were characterized by XRD (Fig. 2A) to identify the crystallographic structure and crystallinity. It can be observed that the precursor can be identified to monoclinic NH₄V₃O₈ (JCPDS no. 01-089-6614, space group: *P21/m*, *a* = 4.9993 Å, *b* = 8.4230 Å, *c* = 7.8490 Å) (Fig. S1†), and the patterns of the mesoporous VO₂ nanowires and the non-mesoporous VO₂ nanowires can both be indexed as monoclinic VO₂ (JCPDS no. 01-081-2392, space group: *C2/m*, *a* = 12.0930 Å, *b* = 3.7021 Å, *c* = 6.4330 Å). Moreover, the XPS spectrum confirmed that the valence of the vanadium is +4 (Fig. S9†). Monoclinic NH₄V₃O₈ consists of the interconnection of the VO₅ square pyramids and VO₆ octahedra within one V₃O₈ layer. The layers are stacked along the *c*-direction with the NH⁴⁺ cations occupying the interlayer space which may be

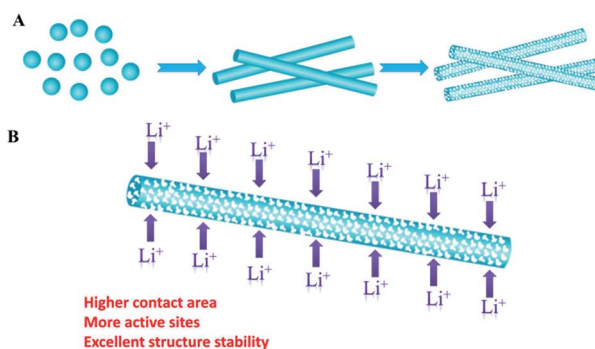


Fig. 1 Schematic illustrations of the fabrication steps and proposed formation mechanism of the mesoporous nanowires (A) and mesoporous nanowires offer high contact area, more active sites and excellent structural stability during lithium insertion/deinsertion (B).

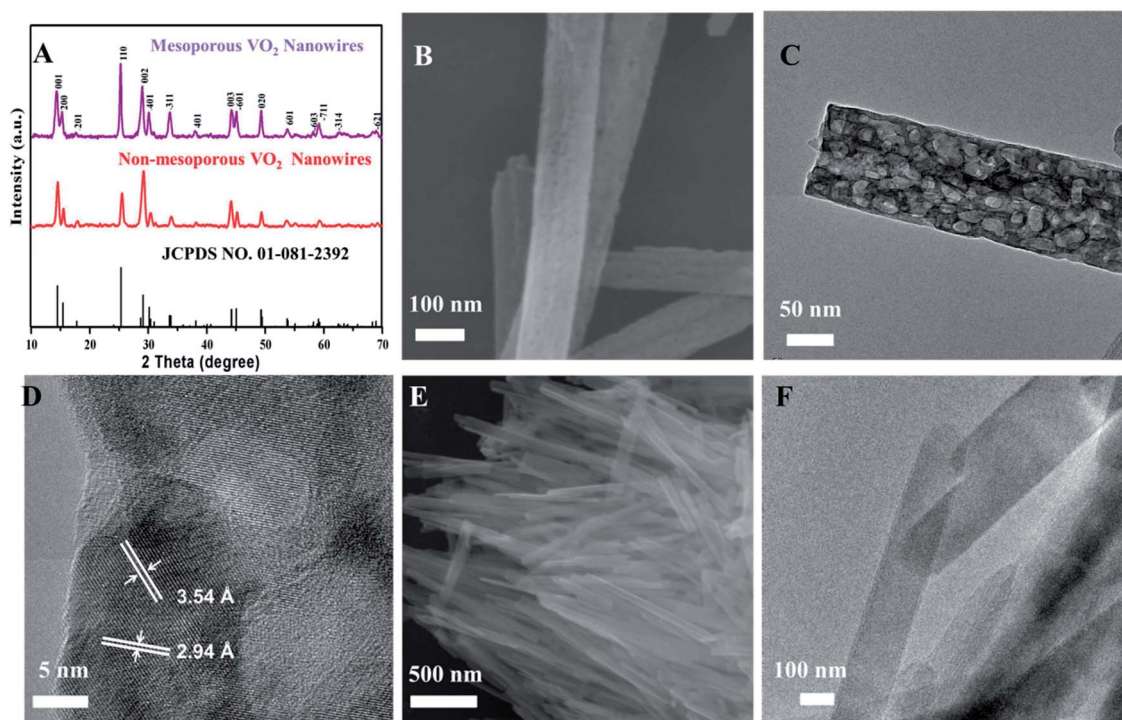


Fig. 2 (A) XRD patterns of non-mesoporous VO₂ nanowires (red), mesoporous VO₂ nanowire (purple); SEM image (B), TEM image (C) and HRTEM image (D) of mesoporous VO₂ nanowire; SEM image (E) and TEM image (F) of non-mesoporous VO₂ nanowires.

suitable for the generation of N₂ and H₂O while retaining the morphology.^{47–49} Monoclinic VO₂ consists of double layers of V₄O₁₀ type.²⁶ According to the *o-t-o* diffusion,⁵⁰ the NH₄⁺ ion occupies a cubic close packed lattice of octahedral sites offering higher energy barrier. Thus, better electrochemical performance of mesoporous VO₂ nanowires is expected than that of NH₄V₃O₈ nanowires (Fig. S2[†]).

SEM and TEM techniques were further applied to characterize the morphology and detailed structure of the as-prepared samples. As shown in the SEM image, the precursor NH₄V₃O₈ nanowires are randomly placed and the diameter of the nanowires is in the range of 80–100 nm (Fig. S3[†]). After annealing, the mesoporous VO₂ nanowires retain the morphology with 120 nm in diameter, indicating that upon the emitting of internally born nitrogen and water, the diameter of the nanowire expands in the annealing process (Fig. 2B). Pores about 20 nm can be easily observed both inside and on the surface of the VO₂ nanowires (Fig. 2C). In order to further identify the pore size and the structure of the mesoporous VO₂ nanowires, HRTEM is carried out (Fig. 2D). The good crystallinity of mesoporous VO₂ nanowires was confirmed by HRTEM. The lattice fringe spacings of 2.94 and 3.53 Å are recognized, corresponding to the (111) and (110) plane of VO₂, respectively, as confirmed by the XRD pattern (Fig. 2A). As shown in Fig. 2E, the non-mesoporous VO₂ nanowires are also 120 nm in diameter. On the surface of the non-mesoporous nanowires, no obvious pore is observed (Fig. 2F). Thus, without undergoing the decomposition process, the non-mesoporous VO₂ nanowires show no signs that pores are on the surface or inside the solid nanowires.

BET is an efficient tool to identify the pore distribution. The N₂ adsorption-desorption isotherms and the corresponding Barrett-Joyner-Halenda (BJH) pore-size distributions of mesoporous VO₂ nanowires are shown in Fig. 3. The mesoporous VO₂ nanowires illustrate typical type-IV N₂ adsorption isotherms with distinct H3 hysteresis loops that can be linked to slit-shaped pores, contributing very high surface area of 46.7 m² g⁻¹ (Fig. 3A), which is over two times as large as the precursor NH₄V₃O₈ nanowires (15.1 m² g⁻¹) (Fig. S4[†]), more than five times as high as non-mesoporous VO₂ nanowires (8.0 m² g⁻¹) (Fig. 3B). The pore size distribution is among 30 nm in diameter. When compared to the NH₄V₃O₈ nanowires and the non-mesoporous VO₂ nanowires, the mesoporous VO₂ nanowires exhibits much higher pore volume (0.6 cm³ g⁻¹), which is five times and ten times higher than those of NH₄V₃O₈ nanowires (0.12 cm³ g⁻¹) and non-mesoporous VO₂ nanowires (0.06 cm³ g⁻¹), respectively, indicating that the highest pore volume may

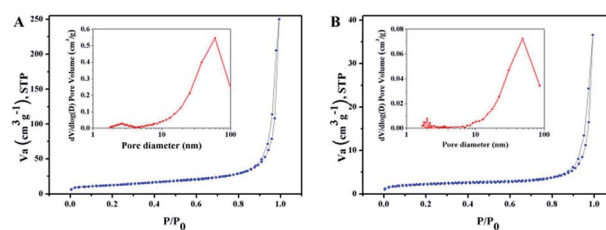


Fig. 3 Nitrogen adsorption and desorption isotherms and pore size distribution (inset) of mesoporous (A) and non-mesoporous (B) VO₂ nanowires.

be attributed to the pores in mesoporous VO₂ nanowires. The resulting large surface area may provide more active sites for lithium intercalation and deintercalation, which may result in better electrochemical performance.

Coin cells of metallic lithium as anode were assembled to investigate the electrochemical performance of as-prepared samples. As shown in the CV curves, a pair of redox peaks occurs (Fig. 4A) when cycled between 2 and 3 V, corresponding to the reaction (2):²⁶



The difference in anodic and cathodic peaks is due to the polarization of the as-prepared samples. In order to evaluate the polarization, the overpotential is calculated from the difference between charge potential and discharge potential at the half reversible capacity, noted as $\Delta V(Q/2)$.²⁶ The overpotential of mesoporous nanowires (0.23 V) is much lower than that of non-mesoporous nanowires (0.36 V), indicating excellent reversibility. The CV curves match well with the charge–discharge voltage profiles (Fig. 4B). And during the first three cycles, the shapes of the charge–discharge curves are almost identical, indicating excellent reversibility of the lithium insertion/deinsertion process. One voltage plateau is clearly observed which is quite consistent with the CV curves. The cycling performance at different current densities ranging from 100 mA g⁻¹ to 1000 mA g⁻¹ is performed. At the current density of 100 mA g⁻¹, the mesoporous VO₂ nanowires deliver initial discharge capacity of 188 mA h g⁻¹ which is a bit higher than that of non-mesoporous nanowires (175 mA h g⁻¹) (Fig. 4C). The capacity of mesoporous VO₂ nanowires witness no obvious decay after 100 cycles with discharge capacity of 180 mA h g⁻¹, corresponding to capacity retention of 95%, exhibiting excellent cycling stability, much higher than that of non-mesoporous VO₂ nanowires (78%). At

current density of 500 mA g⁻¹, the initial capacities of both samples show greater difference (Fig. 4D). The initial discharge capacity of mesoporous VO₂ nanowires (143 mA h g⁻¹) is much higher than that of non-mesoporous nanowires (116 mA h g⁻¹). After 300 cycles, the capacity of mesoporous VO₂ nanowires is 127 mA h g⁻¹, corresponding to 89% of its initial capacity. On the contrast, non-mesoporous VO₂ nanowires fade fast to below 100 mA h g⁻¹ after 100 cycles. All these indicate that the mesoporous VO₂ nanowires show excellent cycling stability.

Further, at high current density of 1000 mA g⁻¹, each sample, during first cycle, delivers a discharge capacity of 105 and 84 mA h g⁻¹, respectively, demonstrating even greater difference in initial capacity between mesoporous and non-mesoporous VO₂ nanowires (Fig. 4E). Mesoporous VO₂ nanowires demonstrate excellent stability with a discharge capacity of 95 mA h g⁻¹, corresponding to 90% of its initial capacity after 500 cycles which is far higher than that of non-mesoporous nanowires (75%). Obviously, the mesoporous VO₂ nanowires exhibit higher capacity and better cycling performance than non-mesoporous VO₂ nanowires at both low and high current density.

The rate capability is a significant evaluation in lithium battery. Thus, both mesoporous and non-mesoporous VO₂ nanowires were cycled at various current density ranging from 100, 200, 500, 1000 mA g⁻¹ and finally back to 100 mA g⁻¹. The mesoporous VO₂ nanowires are with discharge capacities of 186, 165, 141, 101 and 180 mA h g⁻¹ (Fig. 4F), which exhibit higher capacities than those of non-mesoporous VO₂ nanowires (175, 150, 122, 82 and 142 mA h g⁻¹). The capacity retention is higher than those of mesoporous nanowires. All these indicate that even after suffering from rapid change of the current density, the electrode exhibits more stable and higher capacity at each current, indicating a degree of mechanical integrity during lithium ion insertion/deinsertion process. When the

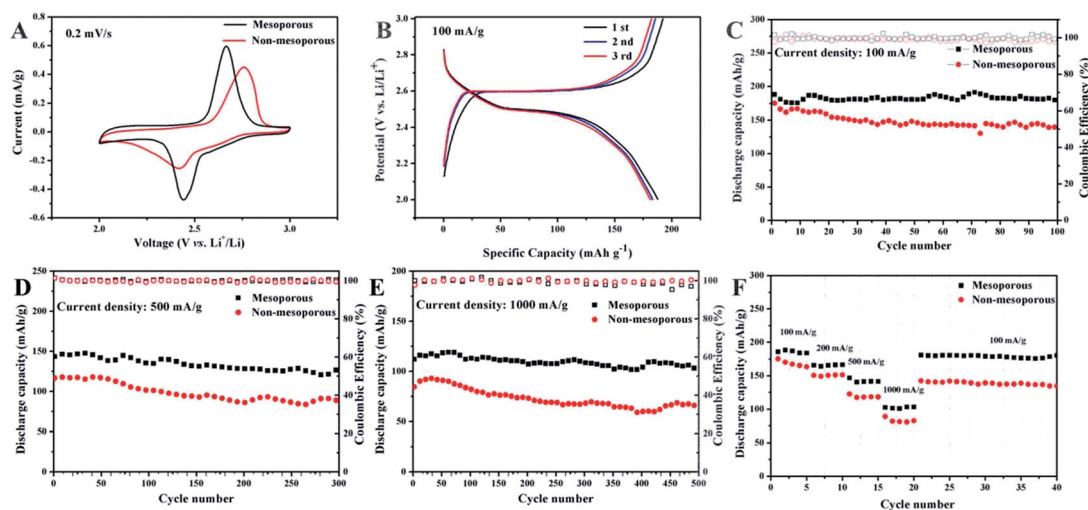


Fig. 4 (A) CV curves of mesoporous and non-mesoporous VO₂ nanowires at scan rate of 0.2 mV s⁻¹; (B) charge and discharge curves of mesoporous VO₂ nanowires during the first three cycles at 100 mA g⁻¹; cycling performance of mesoporous and non-mesoporous VO₂ nanowires at different current densities of 100 mA g⁻¹ (C), 500 mA g⁻¹ (D), and 1000 mA g⁻¹ (E); (F) rate performance of mesoporous and non-mesoporous VO₂ nanowires.

current density returns to 100 mA g^{-1} , about 96% of the initial capacity is recovered and there is no obvious capacity loss during the following 20 cycles. Our results demonstrate the excellent rate capability and structural stability of mesoporous VO_2 nanowires. EIS is a powerful tool to reveal the kinetics of the lithium ion extraction/insertion process which is related to the electron transport and lithium ion diffusion length.⁵¹ EIS of mesoporous and non-mesoporous VO_2 nanowires is carried out. In this circuit, R_s represents the Ohmic resistance of the electrode system, including the electrolyte and the cell components. R_{ct} represents the charge transfer resistance. CPE and Z_w are the double layer capacitance and the Warburg impedance,⁵¹ respectively. All the Nyquist plots are composed of a depressed semicircle in the medium-frequency region followed by a slanted line in the low-frequency region (Fig. S5†). The mesoporous VO_2 nanowires show a much lower resistance of 40Ω when compared with non-mesoporous VO_2 nanowires (68Ω), indicating that the mesopores indeed enhance the charge transfer and modify the kinetics of VO_2 mesoporous nanowires.

To explain the electrochemical results, it is important to consider the one-dimensional mesoporous structure that exerts such a significant influence on enhanced electrochemical performance. Three possible factors may result in the enhanced cycle stability and rate capability. First, the performance is related to the high surface area. The high surface area of mesoporous nanowires ($46.71 \text{ m}^2 \text{ g}^{-1}$), which is 5 times higher than non-mesoporous nanowires, is attributed to the pores in the surface and inside of the nanowires. The hierarchical interconnected pores are able to offer more active sites on the surface or near-surface region for lithium intercalation/deintercalation.¹ Second, according to the diffusion formula $t = L^2/D$ (where t is the diffusion time, L is the diffusion length and D is the diffusion constant),²⁵ the shorter diffusion length will provide less diffusion time, leading to better electrochemical property. The mesopores in the nanowire shorten the Li^+ diffusion distance, strongly increase the lithium mobility and modify the kinetics of the lithium intercalation/deintercalation. Third, this kind of hierarchical structure shows that the mesoporous nanowires restrain the aggregation of nanowires and offer facile stress relaxation to accommodate the volume variations. Thus, this kind of unique pore structure provides more active sites and keeps the surface area larger to make full use of the advantages of nanostructured materials.

Conclusions

Mesoporous VO_2 nanowires have been designed and prepared through facile hydrothermal method followed by annealing process. The thermal decomposition of $\text{NH}_4\text{V}_3\text{O}_8$ to generate nitrogen and water vapor creates to disordered interconnected mesopores. Thus, high specific surface area ($46.7 \text{ m}^2 \text{ g}^{-1}$) is achieved. When evaluated as cathode for lithium batteries, mesoporous VO_2 nanowires exhibits stable cycling performance and excellent rate capability, thus is regarded as a promising candidate for lithium batteries. The capacity of mesoporous VO_2 nanowires witness no obvious decay after 100 cycles with discharge capacity of 180 mA h g^{-1} , corresponding to capacity

retention of 95%, exhibiting much higher capacity and better capacity retention than those of non-mesoporous VO_2 nanowires (174 mA h g^{-1} and 78%), respectively. Mesoporous VO_2 nanowires exhibit a discharge capacity of $104.7 \text{ mA h g}^{-1}$ at the current density of 1000 mA g^{-1} , and 90% of its initial capacity is retained after 500 cycles which is almost two times as high as the capacity of non-mesoporous VO_2 nanowires. This superior performance may benefit from the mesopores, which shorten ion transport route and modify the kinetics during intercalation/deintercalation. Moreover, the disordered interconnected pores are able to offer more active sites for lithium intercalation/deintercalation. Therefore, the facile and effective method can be extended to fabricate nanopores in nanomaterials and constructing this kind of unique one-dimensional mesoporous structure can be utilized for further operation in the fields of energy storage and transportation, such as Li-air battery.

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