



Supporting Information

for *Small*, DOI: 10.1002/sml.201902141

Langmuir–Blodgett Nanowire Devices for In Situ Probing of
Zinc-Ion Batteries

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Supporting Information

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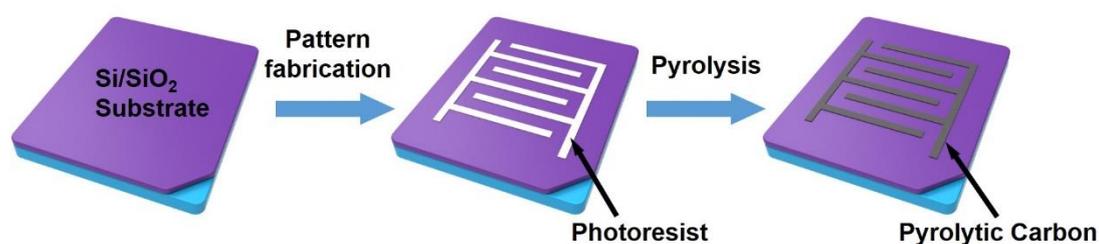
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Figure S1. Schematic illustration of the fabrication processes of pyrolytic carbon current collectors.

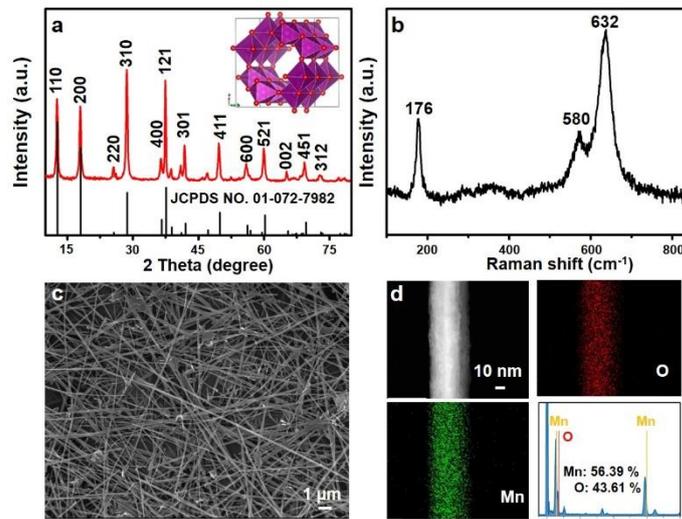


Figure S2. Characterization of MnO₂ nanowires. a) XRD pattern, b) Raman spectroscopy, c) SEM, d) EDS.

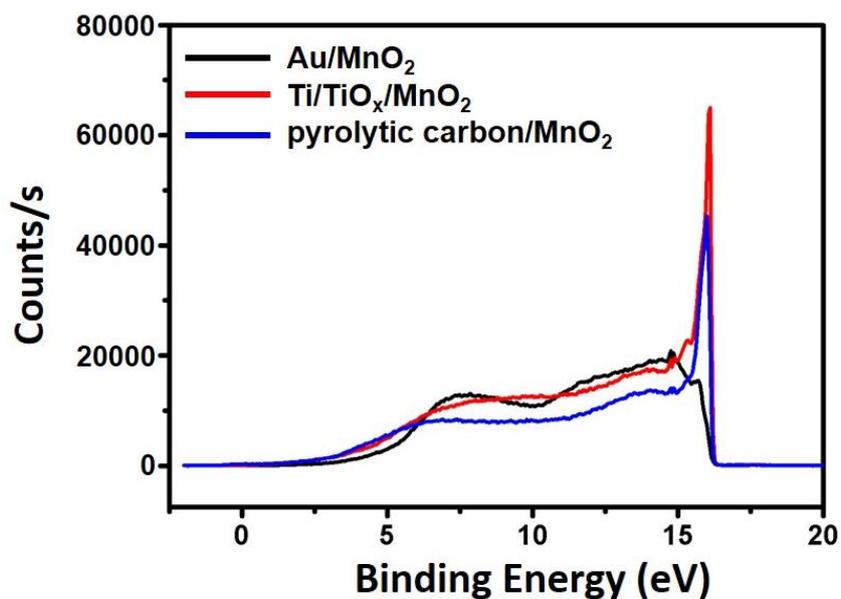


Figure S3. The ultra violet photoelectron spectroscopy of different nanowire/current collector interfaces.

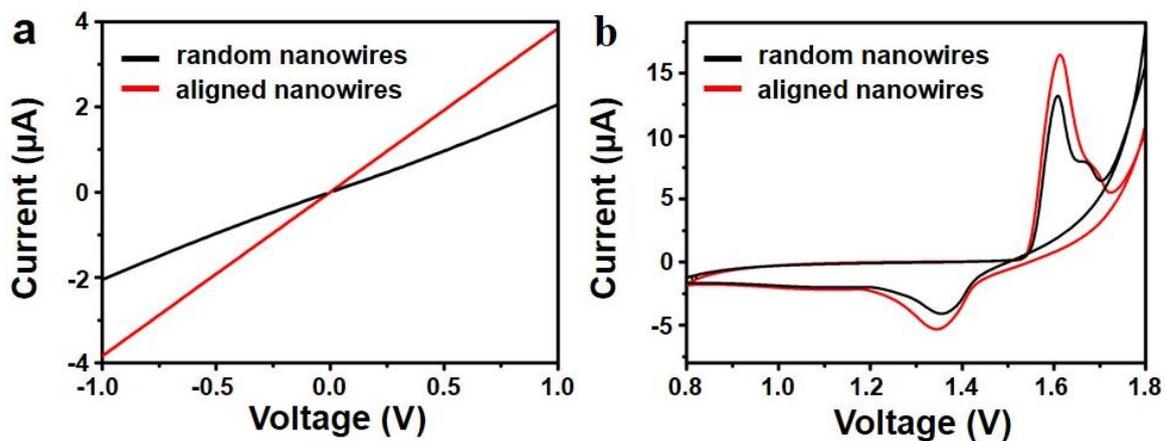


Figure S4. Characterization of LBNW film. a) I-V curves of aligned NWs and random NWs with pyrolytic carbon current collector, respectively. b) CV curves of aligned NWs and random NWs with pyrolytic carbon current collectors in ZnSO_4 electrolyte with pre-added Mn^{2+} .

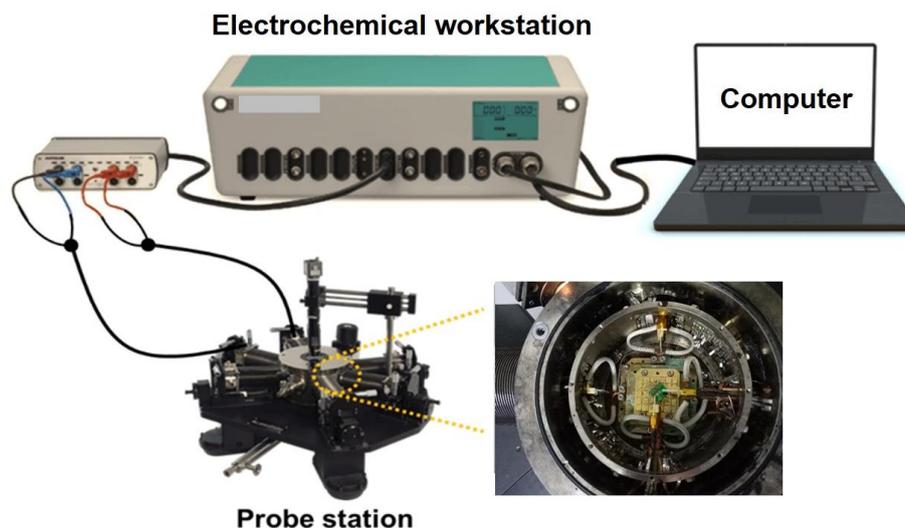


Figure S5. Measurement equipment layout with the three-dimensional view: Electrochemical workstation and probe station combined system.

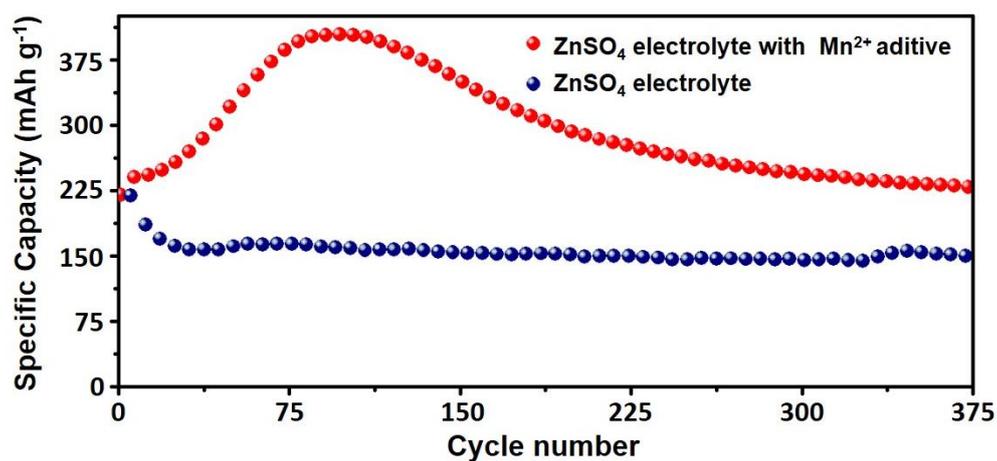


Figure S6. Cycling performance of MnO_2 LBNWs in ZnSO_4 electrolyte and ZnSO_4 electrolyte with pre-added Mn^{2+} , respectively.

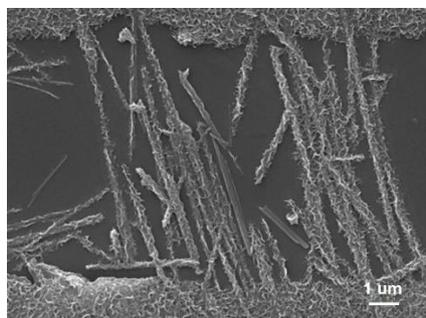


Figure S7. SEM image of MnO_2 nanowires after 100 cycles.

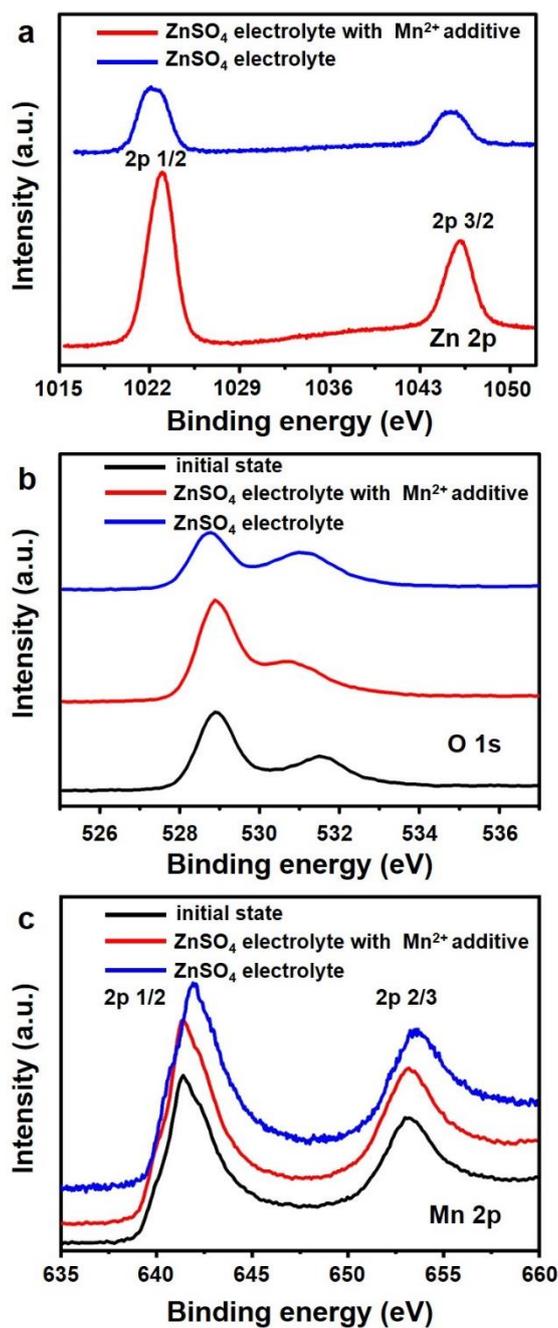


Figure S8. High-resolution XPS characterization of LBNWs at different states. a) Zn 2p, b) O 1s, c) Mn 2p XPS spectroscopy for MnO₂ electrode materials at different states: black, initial state; red, after 100 galvanostatic charging at 1 $\mu\text{A cm}^{-2}$ in ZnSO₄ electrolyte; blue, after 100 galvanostatic charging at 1 $\mu\text{A cm}^{-2}$ in ZnSO₄ electrolyte with MnSO₄ additive.

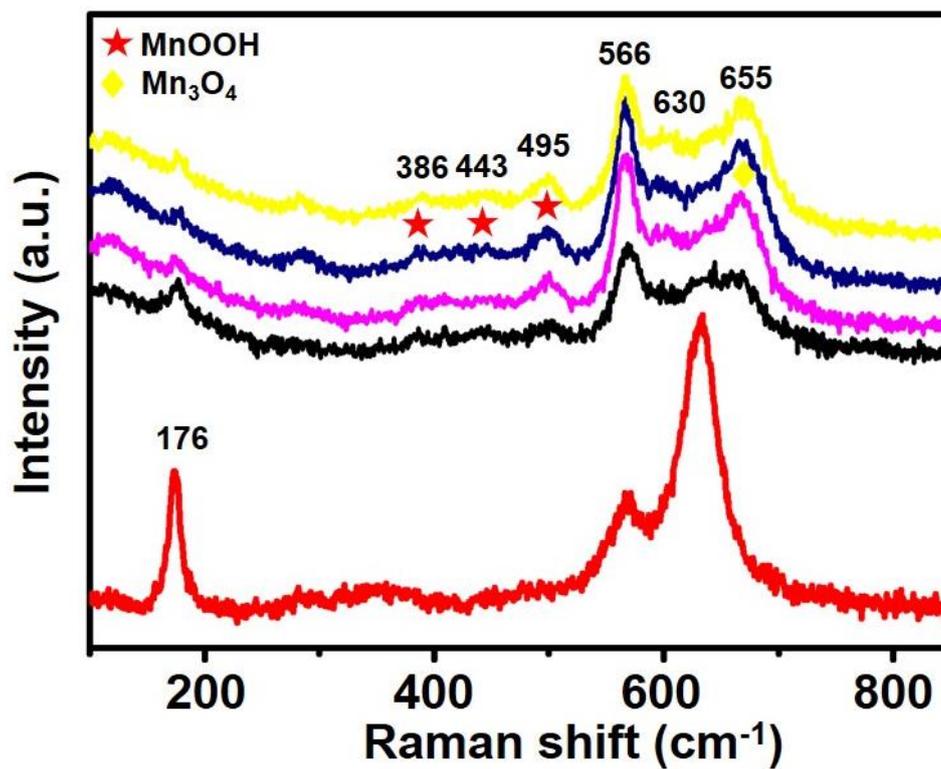


Figure S9. Raman spectroscopy evolution of MnO_2 LNNWs at different charge/discharge states. Red, initial state; dark, after 10th galvanostatic charging; pink, after 20th galvanostatic charging; navy, after 50th galvanostatic charging; yellow, after 375th galvanostatic charging.

Table 1. Zn, Mn, O element ratio after 100 cycles.

Element	Ratio
Oxygen	33.5
Manganese	43.6
Zinc	23.3